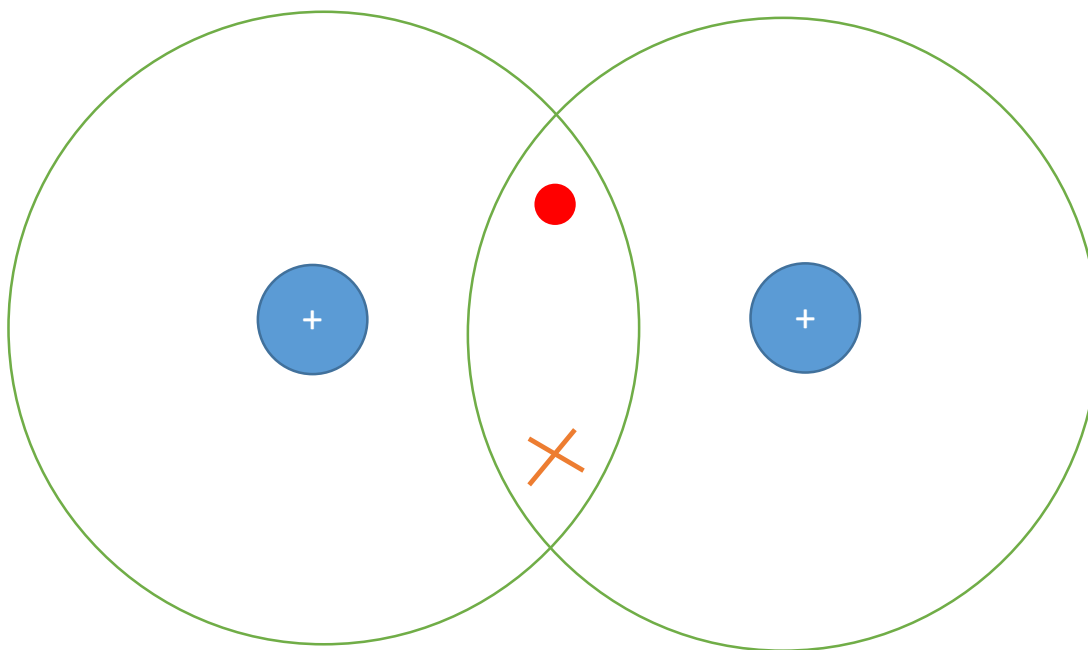


Topic	Covalent bonding and dot and cross diagrams	Level	GCSE (or any other course for students aged 11-16)
Outcomes	<ol style="list-style-type: none"> 1. To understand how a covalent bond is formed 2. To be able to use molecular and displayed formula 3. To draw dot and cross diagrams for simple covalent molecules involving single, double and triple bonds 		

This diagram shows two hydrogen atoms. Annotate the diagram to show all the attractive and repulsive forces that would exist.



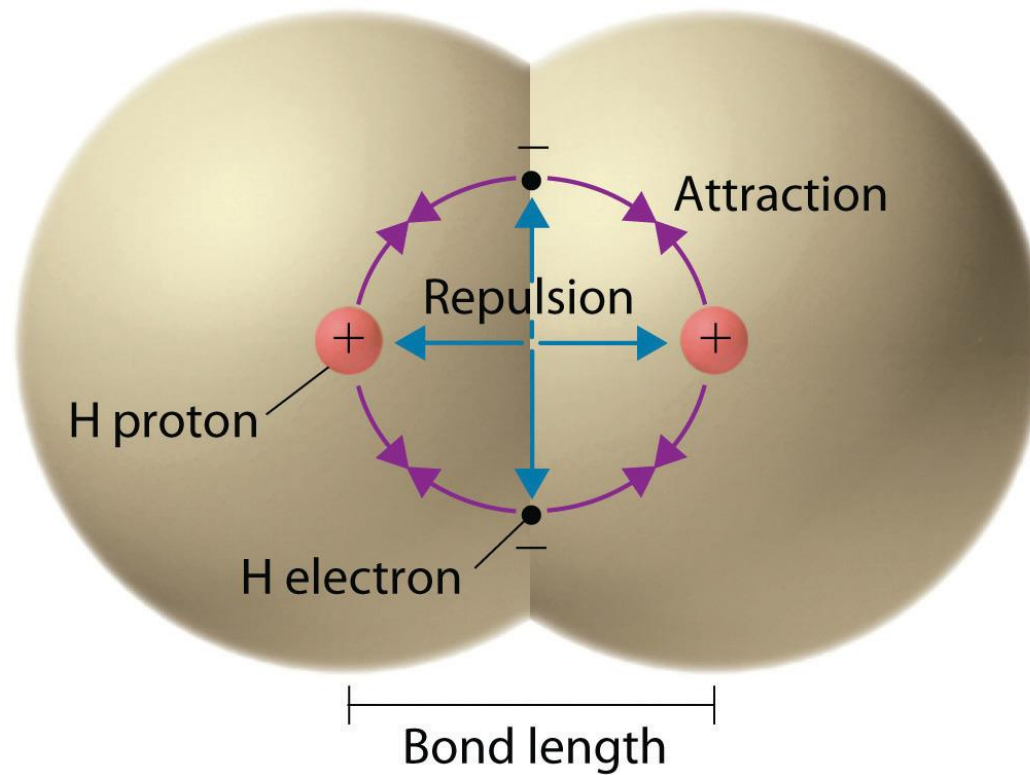
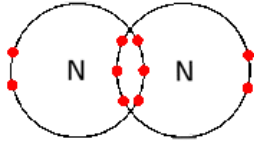


Figure 1 <http://2012books.lardbucket.org/books/principles-of-general-chemistry-v1.0/s12-05-lewis-structures-and-covalent-.html>

Molecular formula	Displayed formula	Dot and cross diagram
H ₂	H-H	
HCl	H-Cl	

NH_3		
	$\text{O}=\text{C}=\text{O}$	
		
C_2H_4		

HCN		
CH ₃ COCl		