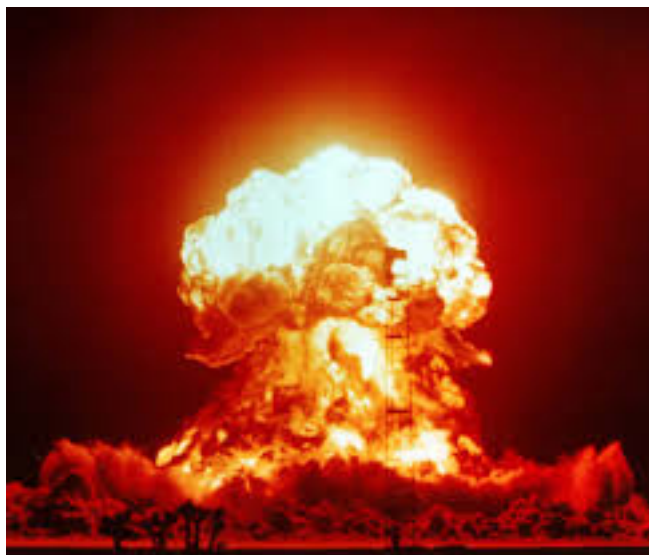


## Rates of Reaction: thinking questions

So, you think you understand rates of reaction?

With your partner, think and discuss these questions below



1. Does the picture above give you any information about the rate of this reaction?
2. What does the term “rate of reaction” mean to you? What are the units?
3. Why don't all chemical reactions occur at the same rate?
4. Do exothermic reactions occur faster than endothermic reactions?
5. Which reaction would you expect to happen faster? Explain.

1.  $A(l) + B(l) \rightarrow AB(l)$
2.  $A(g) + B(g) \rightarrow AB(g)$

[https://www.youtube.com/watch?feature=player\\_embedded&v=TBLr XrooLs](https://www.youtube.com/watch?feature=player_embedded&v=TBLr XrooLs)

6. Does increasing the concentration of a reactant always lead to an increase in the rate of reaction?

Progress: have a look at this resource below to model rates of reaction

<http://www.thescienceteacher.co.uk/resources/How+to+understand+rates+of+reaction+graphs.pdf>